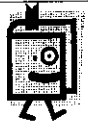


You **MUST copy** the **question if given on the board** and answer it using complete sentences.



Reading Minute Monday

Title: _____

\$2.00 Summary

Directions: With each word worth 10 cents, write a \$2.00 summary of the learning from the lesson. In other words, create a summary that is 20 words.



"Chart"ering New Territory Tuesday

Directions: Create a data table and a line graph using the data provided.

Number of new students enrolled each year at JMS:
2009 (50); 2010 (56); 2011 (65); 2012 (40); 2013 (49)

Title: _____

Title: _____



Teacher's Choice
"Throwback Thursday"

Independent vs. Dependent Variables

Underline the independent variable, and circle the dependent variable.

1. How will the amount of fertilizer used on a field affect the number of earthworms found there?
2. How does the distance from an eye chart affect the number of letters that are recognized on a line?
3. How does the amount of oxygen in the water affect the oyster population?
4. How does the size of a bicycle tire affect the distance it will travel when it is pedaled in a given number of times?



Flocabulary Friday

1. Create 4 notecards for the following words:
 - Element
 - Molecule
 - Compound
 - Mixture
2. Notecards should include the following:
 - Word
 - Picture (colored)
 - Definition
 - Examples (at least 2)

Neatly done

WEDNESDAY WILD WRITING EXPLORATION!

Directions: Answer the following question. Be sure to use the rubric provided to check your writing after you have answered the question.

The tables below show information about two liquids and three solids.

Mass and Volume of Three Solids

Density of Two Liquids

Liquid	Density (g/mL)
A	1.0
B	2.0

Solid	Mass (g)	Volume (mL)
X	20	5
Y	5	10
Z	15	10

Derek was given two colorless liquids and a table listing their densities. He also was given three irregularly shaped solids. He measured the mass and volume of each solid. His results are shown above in the table on the right.

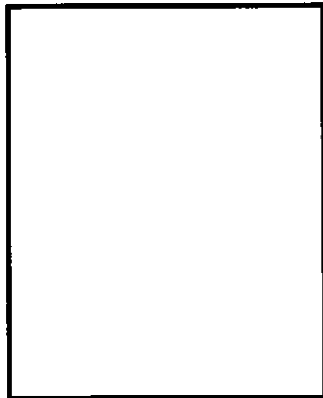
a. Find the densities for solids X, Y, and Z.

Density of X: _____ Density of Y: _____ Density of Z: _____

b. Predict which solid or solids will **sink** in liquid A.

c. Predict which solid or solids will **float** in liquid B.

d. Draw a diagram to represent the three solids and 2 liquids.



e. Explain your diagram. Why did the solids and liquids end up the way you drew them?

	I have read over my response now that I have finished writing it.
	I restated the statement/question in my response (HIGHLIGHT IN GREEN)
	I provided supporting details (HIGHLIGHT IN YELLOW)
	My concluding sentence brings all of the ideas together (HIGHLIGHT IN PINK)
	I have <u>at least 3 sentences minimum</u> in my response.
	I used proper punctuation and capital letters where needed
	I wrote so neat, others can read it
	My response makes sense and is accurate/relevant.

conductivity: a substance's ability to carry an electrical current.

ions: atoms that have a positive or negative electrical charge.

reactive: takes part in a chemical reaction.

malleable: able to be shaped or bent.

Metals combine easily with other metals to form alloys. An alloy is a mixture of one metal with either another metal or a nonmetal. Alloys increase some characteristics of the original metal—usually strength. For example, steel, which is an alloy of iron and carbon, is much stronger than plain iron. Other examples of alloys are brass, which is a mixture of copper and zinc; bronze, which is a mixture of copper and tin; and pewter, which is tin, copper, and lead.

What makes a metallic element different from other elements?

When people describe metal, they usually use words like *hard*, *shiny*, and *solid*. Stomp on an aluminum soda can, though, and it easily crumples beneath your foot. Shine light at a chunk of lead, and the dull, dark surface absorbs most of it. Mercury, however, flows just like any other liquid. Even so, aluminum, lead, and mercury are all metals. In fact, nearly three-quarters of all the elements are considered metals. They aren't all exactly alike, but they do share a few traits.

The first is **conductivity**. The atoms of metallic elements easily lose electrons and become **ions** with positive charges. As you know, electricity is electrons on the move. Metals are good conductors because electrons flow quickly and easily among their atoms.

Metals are also **reactive**, especially with oxygen. Most metals that come into contact with water or air will oxidize, although some do it more quickly than others. Iron oxide, or rust, is the most familiar form of oxidation, but copper, aluminum, zinc, and nickel also have chemical reactions with oxygen over time. Drop sodium into water, though, and watch out. The oxidation is instant, resulting in an explosive release of energy.

The trait all metals share is the way their atoms are bonded together. At the molecular level, metals consist of positive ions held together by a swirling sea of shared electrons. Metallic atoms have strong bonds because they share all these electrons with each other. Metals are reactive because those electrons are readily available to bind with negative ions of other elements needing electrons. Electrons moving freely among the atoms are also why metals are such good conductors.

Sharing electrons makes metals **malleable**. You can easily bend, mold, or stretch a piece of heated metal without it breaking apart because metallic bonds are strong but not stiff. The atoms slide and stretch without pulling apart because the shared electrons slide and stretch with them.

Metals can be divided into several different categories. For example, the alkali metals, like sodium, potassium, and lithium, are extremely reactive. The transition metals are by far the largest category. Gold, silver, and copper are examples of transition metals.

Periodic Table of the Elements

The periodic table displays elements in groups and periods. Key groups include the alkali metals (Group 1), transition metals (Groups 3-10), halogens (Group 17), and noble gases (Group 18). The lanthanide and actinide series are shown at the bottom.